VR Science Lab

Hydrogen Peroxide with Potassium Iodide

Affectionately known as elephant toothpaste, the chemical reaction between the peroxide and potassium iodide shoots out a column of foam. If you add food coloring, you can customize the "toothpaste" for holiday-colored themes

Hydrogen peroxide is a <u>chemical compound</u> with the formula H_2O_2. In its pure form, it is a very pale blue <u>liquid</u>, slightly more <u>viscous</u> than <u>water</u>. It is used as an <u>oxidizer</u>, <u>bleaching</u> agent, and <u>antiseptic</u>, usually as a dilute solution (3–6% by weight) in water for consumer use, and in higher concentrations for industrial use. Concentrated hydrogen peroxide, or "<u>high-test peroxide</u>", decomposes explosively when heated and has been used as a <u>propellant</u> in <u>rocketry</u>.

Hydrogen peroxide is a <u>reactive oxygen species</u> and the simplest <u>peroxide</u>, a compound having an oxygen–oxygen <u>single bond</u>. It decomposes slowly when exposed to light, and rapidly in the presence of organic or reactive compounds. It is typically stored with a <u>stabilizer</u> in a weakly acidic solution in a dark bottle to block light. Hydrogen peroxide is found in biological systems including the human body. Enzymes that use or decompose hydrogen peroxide are classified as <u>peroxidases</u>.

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